

# SPECIAL

## OXYGENIZED DRINKING WATER



**SPECIAL** Water is made by

**SPECIAL Filter** dissolving a high concentration of OXYGEN in Natural Tap Water using the Mineral known as

### Temporary Water Hardness

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## What is **SPECIAL** Water ?

The term **SPECIAL Water** is just to Label used to describe ideal drinking water as provided by Nature



All **NATURAL Water** Consisting of H<sub>2</sub>O molecules and OH<sup>-</sup> and H<sup>+</sup> ions in very small quantities. These are ions that give acidity and alkalinity to water. When these ions are equal in number, the water is **NATURAL** and **NEUTRAL**. When any other Ion whether its **CATION** or **ANION** is added to water its not natural anymore.

H<sup>+</sup> ions addition makes the water **ACIDIC** while OH<sup>-</sup> ions make the water more **ALKALINE**.

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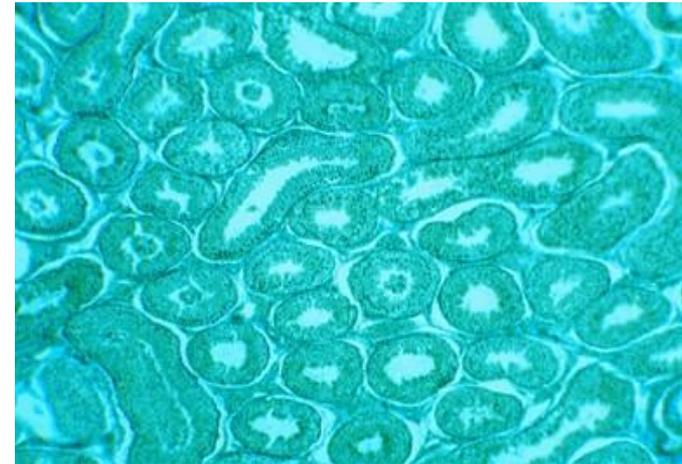
# What is SPECIAL Water ?

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## **NOW TRY TO UNDERSTAND...**

OH<sup>-</sup> Water has more **Oxygen** and alkaline minerals (Calcium and Magnesium) described as CaCO<sub>3</sub> equivalent than Natural drinking water or **ACIDIC** and **SODIUM** enriched tap water, that's been treated with **ION EXCHANGE** Process.

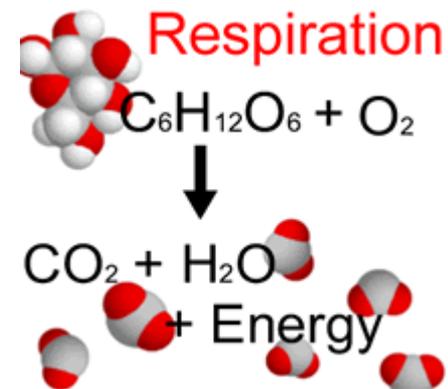
**SPECIAL Water** is high oxygen content in a stable bias OH<sup>-</sup> form that is needed by all human beings cells. Each of these ions are bonded with an alkaline Calcium Carbonate minerals in very beneficial colloided form.



# Oxygen in SPECIAL Water

Increase in free carbon dioxide (CO<sub>2</sub>) is often concomitant with decrease in **Oxygen**. In **Natural** waters as CO<sub>2</sub> as an end product and leaves water, **Oxygen** content raises rapidly and at pH of 7.4 it would go up due to the release of CO<sub>2</sub> out of the water and diffusion of Oxygen will take place. The **Oxygen** loading capacity is increased at low CO<sub>2</sub> level in water.

Drinking **Oxygen** enriched water will be shifting oxygen to blood. Drinking high Oxygenated water will effect both metabolism and energy release and kill all depression factors related to acidic water intake.



# pH in SPECIAL WATER

As you have already read the relation of interaction of pH and  $\text{CO}_2$ , it is the change in concentration of  $\text{H}^+$  ions (coupled with  $\text{HCO}_3^-$  and unionized  $\text{H}_2\text{CO}_3$ ) which decides the level of pH of WATER/BLOOD.

## As WELL KNOWN FACT

pH is the negative logarithm of hydrogen ion ( $\text{H}^+$ ) concentration within the pH range of **0 to 14**. Every human being lives the most healthiest life at pH between 7.5 and 9, and body prefers slightly alkaline water, Beverages, Coffee and food cooked in water which is close to neutral pH. This can **only** be done by not exchanging ions (*unlike Ion-Exchange Resins*) from Calcium and Magnesium against Hydrogen or Sodium.

# Alkalinity and Water Hardness

Alkalinity and Water Hardness are closely related. The part of total hardness chemically equivalent to total alkalinity is termed as

**“CARBONATE HARDNESS”**

Which is also referred as

**“TEMPORARY WATER HARDNESS”**

As the Carbonates Hardness are shifted from water (Scale Prevention) with



**CaCO<sub>3</sub>** Crystals that cannot form scale

**H<sub>2</sub>O** Becomes Neutral with keeping it's alkalinity

**CO<sub>2</sub>** Leaves water increasing the diffusion of Oxygen into the water.

# Permanent Water Hardness

Non-carbonate water hardness which is called „*Permanent Water Hardness*“ as it cannot be removed by Water-Softening or Scale Prevention system.

Permanent Water Hardness is

Associated with

**Sulfates** (SO<sub>4</sub>), **Chlorides** (Cl),  
**Silicates** (SiO<sub>3</sub>), **Nitrates** (NO<sub>3</sub>) and **Phosphates** (PO<sub>4</sub>)  
and **NOT** with bicarbonates (HCO<sub>3</sub>) and carbonates (CO<sub>3</sub>)

*More on PERMANENT WATER HARDNESS: Coming Soon*

*Thanks For Reading*